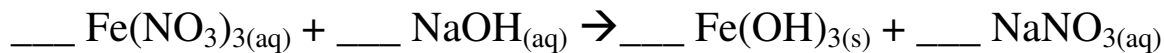


## HW - LIMITING REACTANTS II

NAME \_\_\_\_\_



For each of the following questions find:

- a) What is the limiting reactant? What is the excess reactant?
  - b) How many grams of each product do you expect?
  - c) How many grams of the excess reactant will be left over?
- 1) You begin with 0.354 moles of iron(III) nitrate and 0.562 moles of sodium hydroxide:

- 2) You begin with 25.0g of  $\text{Fe(NO}_3\text{)}_3$  and 35.0g of NaOH:

## HW - LIMITING REACTANTS II

NAME \_\_\_\_\_

- 3) You begin with 403g of iron(III) nitrate and 298g of sodium hydroxide.