HW - LIMITING REACTANTS I

NAME_____

____ $-H_2O_{(1)} + \underline{\qquad} P_2O_{5(s)} \rightarrow \underline{\qquad} H_3PO_{4(aq)}$

For each of the following questions find:

- a) What is the limiting reactant? What is the excess reactant?
- b) How many grams of phosphoric acid do you expect?
- c) How many grams of the excess reactant will be left over?

1) You begin with 3.50 moles of water and 2.00 moles of P_2O_5 :

2) You begin with 23.0g of water and 23.0g of P_2O_5 :

3) You begin with 10.0g of water and 2.00g of P_2O_5 :