NAME

$$
\ldots \ldots \mathrm{H}_{2} \mathrm{O}_{(\mathrm{l})}+\ldots \ldots \mathrm{P}_{2} \mathrm{O}_{5(\mathrm{~s})} \rightarrow \ldots \mathrm{H}_{3} \mathrm{PO}_{4(\mathrm{aq})}
$$

For each of the following questions find:
a) What is the limiting reactant? What is the excess reactant?
b) How many grams of phosphoric acid do you expect?
c) How many grams of the excess reactant will be left over?

1) You begin with 3.50 moles of water and 2.00 moles of $\mathrm{P}_{2} \mathrm{O}_{5}$ :
2) You begin with 23.0 g of water and 23.0 g of $\mathrm{P}_{2} \mathrm{O}_{5}$ :
3) You begin with 10.0 g of water and 2.00 g of $\mathrm{P}_{2} \mathrm{O}_{5}$ :
