**Experiment # 10 – Predicting & Balancing Chemical Reactions**

Part 1 – Predict the products and use your solubility rules chart to determine whether or not a reaction is likely to take place. Write your predictions and the balanced chemical equations as your hypothesis. *This portion is your pre lab and must be completed prior to the lab. If this is not complete then you will not be able to start lab with the rest of the class. If you do not complete the lab because you did not complete the pre lab then your grade will be penalized.*

1. KI + Pb(NO3)2
2. BaCl2 + Pb(NO3)2
3. BaCl2 + K2SO4
4. HCl + NaOH
5. HCl + Na2CO3
6. KCl + Pb(NO3)2
7. BaCl2 + Na2CrO4
8. NaC2H3O2 + NaOH
9. Cu(NO3)2 + NaOH
10. NaNO3 + KCl

Part 2 – Procedure

1. Predict whether or not a double displacement reactions will occur and what will you observe (i.e., a precipitate). 🡪 This is your pre lab activity.
2. Mix together three drops of each solution in a well plate
   1. You will need to keep track of which reaction is in which plate because different reactions have different disposal requirements.
3. Record your observations in the data sheet
   1. Observations should consist of a 4 column table: 1st column for the reaction number 2nd column for the solutions 3rd column for the prediction (hypothesis for each) 4th column for observations and data, you may want to leave additional space for multiple observations.
4. Follow Mrs. Fergusson’s instructions for the disposal.